

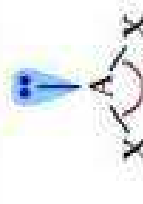


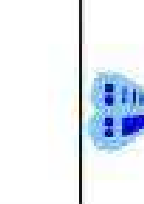









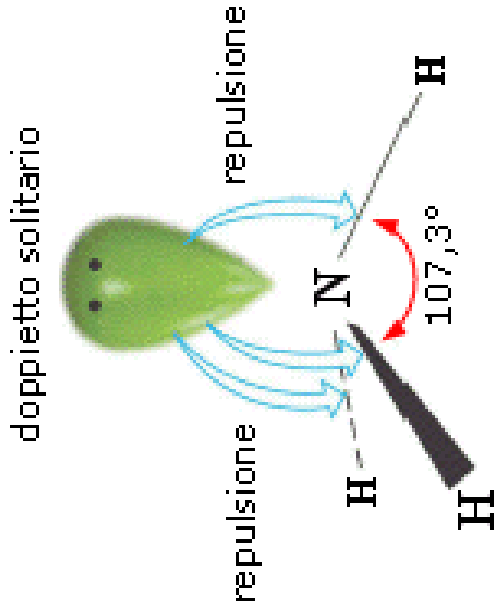
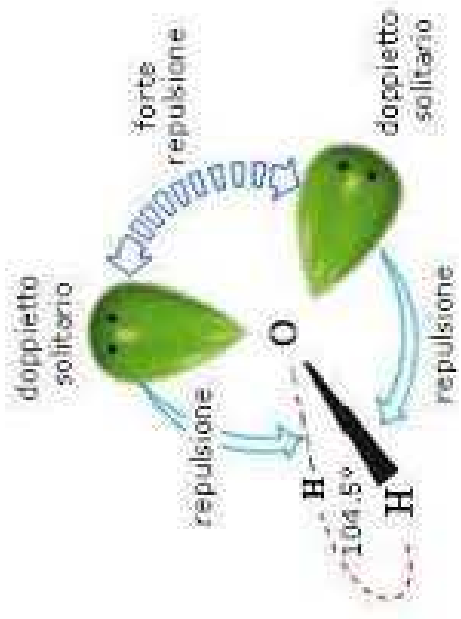
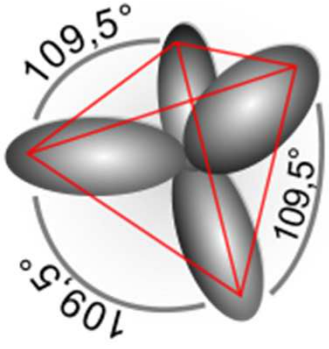
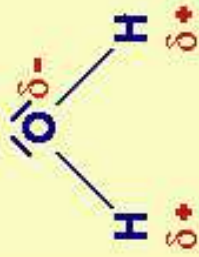


Geometrie VSEPR					
Coppie solitarie					
	0	1	2	3	4
NS=2	 <p>AX<sub>2</sub> lineare 180°</p>				
NS=3	 <p>AX<sub>3</sub> Trigonale planare 120°</p>	 <p>AX<sub>2</sub>E<sub>1</sub> Angolata 120°</p>			
NS=4	 <p>AX<sub>4</sub> Tetraedrica 109° 28'</p>	 <p>AX<sub>5</sub>E<sub>1</sub> Piramidale trigonale 109°</p>	 <p>AX<sub>3</sub>E<sub>2</sub> Angolata 109°</p>		
NS=5	 <p>AX<sub>5</sub> Bipiramidale trigonale 90° 120°</p>	 <p>AX<sub>4</sub>E<sub>1</sub> Alcornocone o cavalletto (Sawhorse) 90° 120°</p>	 <p>AX<sub>3</sub>E<sub>2</sub> a forma di T 90°</p>	 <p>AX<sub>2</sub>E<sub>3</sub> Lineare 180°</p>	
NS=6	 <p>AX<sub>6</sub> Ottaedrica 90°</p>	 <p>AX<sub>5</sub>E<sub>1</sub> Piramidale quadrata 90°</p>	 <p>AX<sub>4</sub>E<sub>2</sub> Planare quadrata 90°</p>	 <p>AX<sub>3</sub>E<sub>3</sub> a forma di T 90°</p>	 <p>AX<sub>2</sub>E<sub>4</sub> Lineare 180°</p>

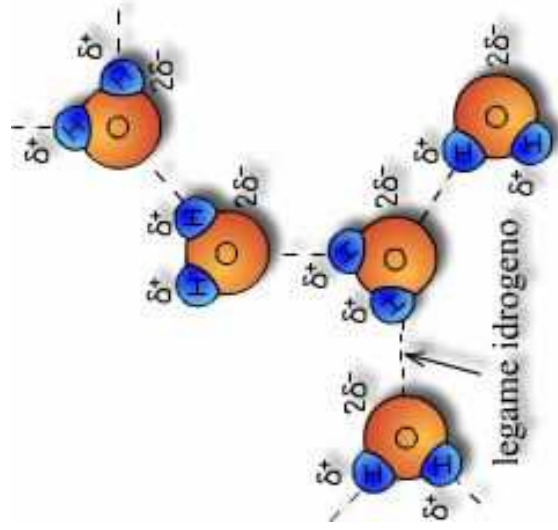




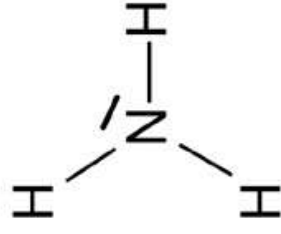
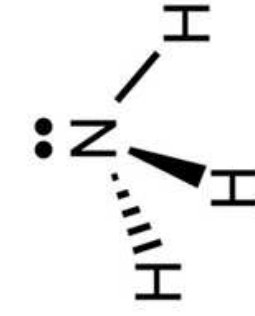
CO<sub>2</sub> è apolare e lineare: i due momenti polari si compensano



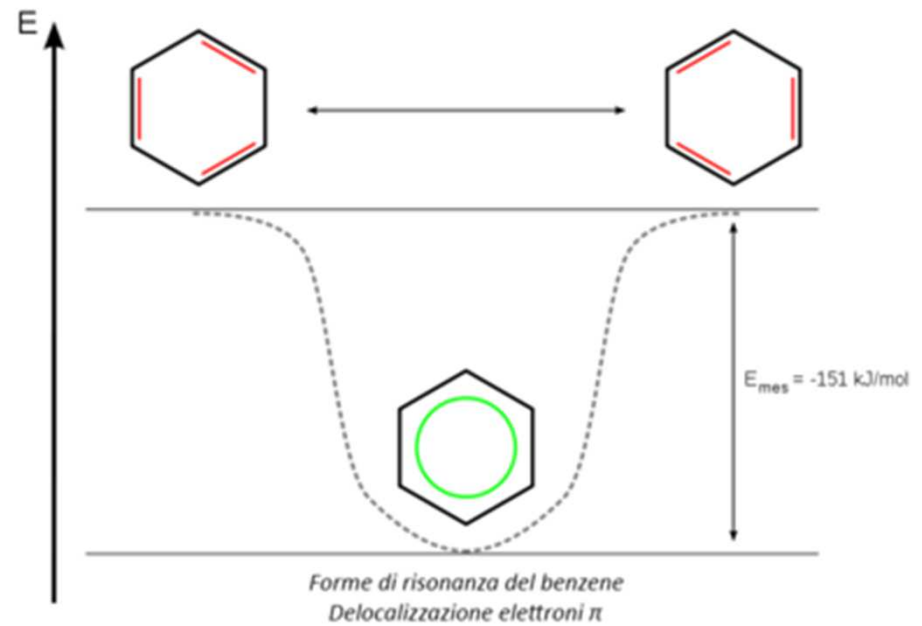
H<sub>2</sub>O è polare, perciò deve avere struttura angolare: i due momenti polari non si compensano



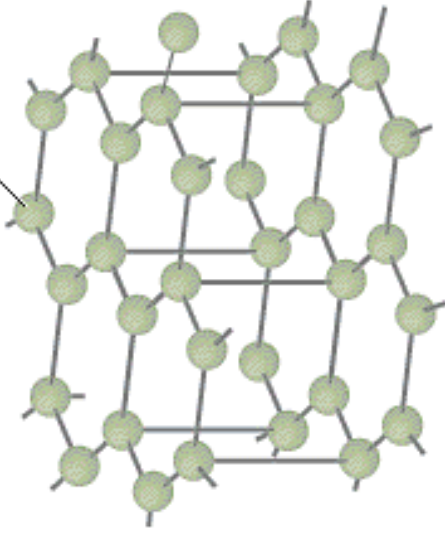
### RAPPRESENTAZIONI DELL'AMMONIACA



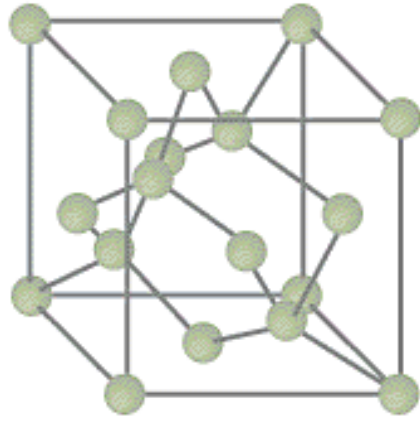
# RISONANZA



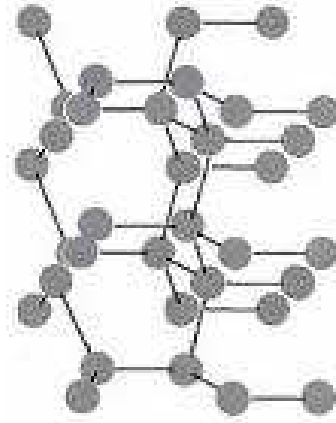
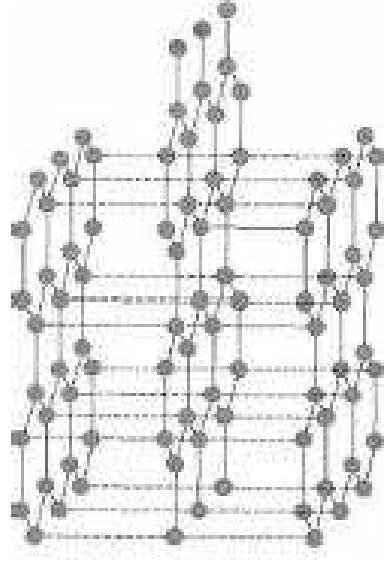
Atomo di carbonio



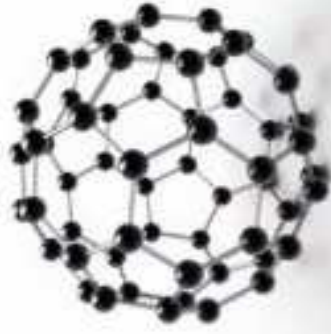
**Grafite**



**Diamante**



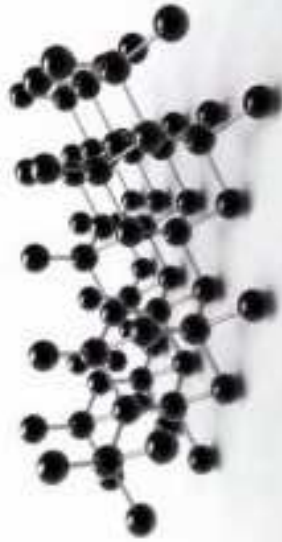
**Figura 1. Diamante e grafite**



fullerene



nanotubo



diamante



grafene